Predict and balance the following reactions. Use the abbreviations (s), (l), (g) and (aq) for the reactants and products. All reactants are aqueous unless otherwise stated. Then, state the overall and net ionic equations.

1. Silver nitrate and sodium sulfate
2. Barium chloride and zinc sulfate
3. Ammonium sulfide is reacted with hydrochloric acid (HCl)
4. Sodium nitrate is reacted with hydrochloric acid
5. Ammonium carbonate and calcium chloride
6. Ammonium bromide and sodium hydroxide
7. Ammonium sulfate and potassium hydroxide are mixed together
8. Potassium phosphate and strontium nitrate
9. Potassium sulfite is reacted with hydrobromic acid (HBr)
10. Magnesium nitrate and sodium hydroxide
11. Equal numbers of moles of potassium hydroxide and phosphoric acid (H₃PO₄) react